a) Low T
(small molecules flying higher, fewer collisions at relatively low speeds)

These slam into the walls of the container not so hard. \( \Rightarrow \) low pressure.

b) There are a few ways to look at this...

Molecules don't have as far to go before hitting wall again. \( \Rightarrow \) more collisions per sec. \( \Rightarrow \) higher pressure.

OR Molecules are closer together, so more collisions per unit area per unit time. \( \Rightarrow \) higher pressure.